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Note : # Neutrons = Mass Number –

Atomic Number. # Neutrons of

Nitrogen-14 = $14 - 7 = 7$. # Neutrons of

Nitrogen-13 = $13 - 7 = 6$. c) The number

of electrons in a neutral atom always

equals the atomic number. The atomic

number of both isotopes is 7, therefore the

number of electrons in both isotopes is

also 7.

Organic Chemistry, 5th Edition Chapter 1

- Section 1.1 ...

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