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Questions pH \u0026 \u0026

| pH value of

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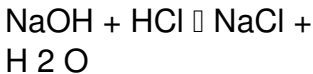
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Wonders Acids Bases  
And Salts Answers

Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization

reaction. Examples:



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Questions and

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Solutions. Acids

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## Acids Bases

Questions Very Short  
Answer Type.

Question 1. Two  
solutions A and B  
have pH values of 5  
and 8 respectively.  
Which solution will be  
basic in nature ...

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Questions with ...~~

Ans:- Salts are  
classified into four

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## Acids Bases

categories on the basis of the strength of acid and base. they are: Salt of strong acid and strong base. e.g NaCl, KCl; Salt of strong acid and weak base. e.g.  $\text{NH}_4\text{Cl}$ ,  $\text{NH}_4\text{I}$ ; Salt of weak acid and strong base. e.g.  $\text{CH}_3\text{COONa}$ ,  $\text{COO}^- \text{Na}^+$  + Salt of weak acid and weak base. e.g.



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$\text{CH}_3\text{COO}^- \text{NH}_4^+$

How is common salt  
found in nature?

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2. Bases react with .  
acids to produce salts  
and water; water to  
produce acids and  
salts; salts to produce  
acids and water;  
neither acids, salts,  
nor water

~~Acids Bases and~~

*Page 19/42*

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~~Salts Multiple Choice  
Questions and  
Answers~~

The salts of strong acids and weak bases give acidic solution having pH less than 7. Example,  $\text{NH}_4\text{Cl}$ , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH

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Acids Bases

And Salts

more than 7.  
Example,  $\text{Na}_2\text{CO}_3$ ,  
Sodium Carbonate  
will have pH more  
than 7. Question 7.

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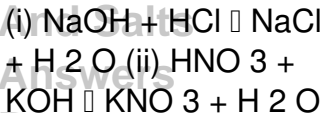
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Answer A reaction in  
which an acid and  
base react with each  
other to give a salt  
and water is termed  
as neutralization  
reaction. For Example:

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## Acids Bases



15. Give two

important uses of washing soda and baking soda. Answer Two important uses of washing soda are:

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Answer. (a) HCl is



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## Acids Bases

acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl. It may be white or light blue or yellow depending upon

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impurities present in  
it.

Answers

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## Acids Bases

### ~~Salts Note Taking~~

### ~~Answers~~

(A) inorganic acid (B)  
organic acid (C) soft  
organic acid (D)  
strong inorganic acid.

Answer. Answer: (a)

5. When milk of  
magnesia reacts with  
acetic acid it produces  
(A) basic salts (B)  
acidic salts (C) neutral  
salt (D) complex salt.

Answer. Answer: (c)

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6. Which of the following phenomena will occur when a small amount of acid is added to water ...

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Turmeric is a natural



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## Acids Bases

indicator. On adding its paste to acid and base separately, which colours would be observed. (a)

Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid and yellow in base (d) Red in acid and blue in base. Answer.

Answer: (b) Yellow in acid and red in base

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Acids Bases

~~Bases~~ Salts

Acids, Bases and Salts In this first quiz on GCSE Chemistry we take a look at one vital part of the subject - acids, alkalis, bases and salts. In it we find out the difference between alkalis and bases, how acids and alkalis react with one another, how

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Acids Bases

neutralisation can produce salts and the relevance of pH values.

Chemistry

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Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste

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Acids react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

~~Acids, bases and salts. 7th Grade Science Worksheets~~



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Acid + Base  $\rightarrow$  Salt  
+ Water. Question 3:  
Rain becomes acidic  
due to. carbon



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## Acids Bases

dioxide. sulphur dioxide. nitrogen dioxide. all of the above. Correct Option is : 4. Solution : Rain becomes acidic because of carbon dioxide, sulphur dioxide and nitrogen dioxide. They dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid ...

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